

Section: Substances Are Made of Atoms

In the space provided, write the letter of the term or phrase that best answers the question.

1. According to the law of definite proportions, any two samples of water, H_2O ,
 a. will be made up only of particles consisting of two atoms of hydrogen and one atom of oxygen.
 b. will have the same ratio of mass of hydrogen to mass of oxygen.
 c. will have a 2:1 ratio of mass of hydrogen to mass of oxygen.
 d. Both (a) and (b)
2. The law that states that mass cannot be created or destroyed in ordinary chemical and physical changes is known as the law of
 a. conservation of mass.
 b. mass action.
 c. multiple proportions.
 d. definite composition.
3. "When two elements combine to form two or more compounds, the mass of one element that combines with a given mass of the other element is in the ratio of small whole numbers." This statement is known as the law of
 a. conservation of mass.
 b. mass action.
 c. multiple proportions.
 d. definite composition.
4. Which two compounds are examples of the law of multiple proportions?
 a. $FeCl_3$ and $Fe_2(SO_4)_3$
 b. O_2 and O_3
 c. CO and CO_2
 d. $FeCl_2$ and $Fe(NO_3)_2$
5. The law of multiple proportions can be partly explained by the idea that
 a. elements can combine in only one way to form compounds.
 b. whole atoms of the same two elements combine to form different compounds.
 c. elements in a compound always occur in a 1:1 ratio.
 d. only atoms of the same element can combine.

Quiz

Assessment

Name _____

Class _____

Date _____